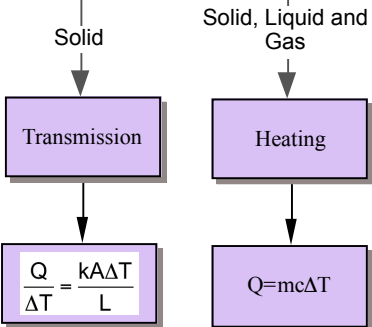


Thermodynamics
Enduring Understanding
 The fluid interactions and thermodynamics process are governed by conservation laws.
Essential Questions
 5. Why is there not a delta in front of symbol Q for heat?
 6. When is work done by a gas undergoing a thermodynamic process?

Energy

Non Mechanical Energy

Heat



Laws of Thermodynamics
Zeroth Law - $T_A = T_B; T_B = T_C; T_A = T_C$
First Law - $\Delta U = Q + W$
Second Law - Heat will not flow spontaneously from a cold object to a hot object. Any system which is free of external influences becomes more disordered (entropy increases) with time.
Third Law - It is impossible to reduce the temperature of a system to absolute zero in a finite number of steps

$$PV = nRT$$

$$\Delta U = Q + W$$

$$K = \frac{3}{2}k_B T$$

Useful Conversions
 1 atm = 101300 Pa
 1000 L = 1 m³



Definitions
 +W = gas is compressed
 -W = gas expands
 +Q = gas gains heat
 -Q = gas loses heat

Graphs
 Work = area beneath P vs. V graph
 Clockwise = -W; Heat Engine
 Counterclockwise = +W; Refrigerator
 Adiabatic curve is STEEPER than Isothermal

